


Binary ionic compounds names and formulas worksheet answers

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Naming Ionic Compounds Practice Worksheet

Name the following ionic compounds:

- 1) NH_4Cl _____
- 2) $\text{Fe}(\text{NO}_3)_3$ _____
- 3) TiBr_3 _____
- 4) Cu_3P _____
- 5) SnSe_2 _____
- 6) GaAs _____
- 7) $\text{Pb}(\text{SO}_4)_2$ _____
- 8) $\text{Be}(\text{HCO}_3)_2$ _____
- 9) $\text{Mn}_2(\text{SO}_3)_3$ _____
- 10) $\text{Al}(\text{CN})_3$ _____

Write the formulas for the following compounds:

- 11) chromium (VI) phosphate _____
- 12) vanadium (IV) carbonate _____
- 13) tin (II) nitrite _____
- 14) cobalt (III) oxide _____
- 15) titanium (II) acetate _____
- 16) vanadium (V) sulfide _____
- 17) chromium (III) hydroxide _____
- 18) lithium iodide _____
- 19) lead (II) nitride _____
- 20) silver bromide _____

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3 • Molecules and Compounds

MOLAR MASS & % COMPOSITION

I. Molar Masses

Given a periodic table, you should be able to calculate the molecular mass (in u's) or the molar mass (in grams) for any element or compound.

Examples: (give answers to two decimal places)

H_2SO_4 98.09 amu	Cl_2 70.90 amu	$\text{Ca}(\text{OH})_2$ 74.10 amu	$\text{HC}_2\text{H}_3\text{O}_2$ 60.06 amu
CO_2 44.01 amu	N_2O 44.02 amu	NaOCl 74.44 amu	Al_2S_3 150.17 amu

II. Fraction and Percent Composition

It is useful to determine how much of a compound's mass is made up of each element. Water, H_2O , for example has a molar mass of 18.02 g. The H's mass is $2(1.0079) = 2.02$ g. The O's mass is 16.00 g.

We can set up fractions for each element: $\text{H} = \frac{2.02}{18.02} = 0.112 = 11.2\%$ $\text{O} = \frac{16.00}{18.02} = 0.888 = 88.8\%$

This is called the **percent composition**. The fraction composition is a good in-between step.

Determine the fraction and percent composition of each element below (answer to one decimal place):

1. H_2SO_4	$\text{H} = \frac{2.02}{98.09} = 2.1\%$	$\text{S} = \frac{32.07}{98.09} = 32.7\%$	$\text{O} = \frac{64.00}{98.09} = 65.2\%$
2. $\text{Ca}(\text{OH})_2$	$\text{Ca} = \frac{40.08}{74.10} = 54.1\%$	$\text{O} = \frac{32.00}{74.10} = 43.2\%$	$\text{H} = \frac{2.02}{74.10} = 2.7\%$
3. $\text{HC}_2\text{H}_3\text{O}_2$	$\text{H} = \frac{4.04}{60.06} = 6.7\%$	$\text{C} = \frac{24.02}{60.06} = 40.0\%$	$\text{O} = \frac{32.00}{60.06} = 53.3\%$
4. CO_2	$\text{C} = \frac{12.01}{44.01} = 27.3\%$	$\text{O} = \frac{32.00}{44.01} = 72.7\%$	
5. N_2O	$\text{N} = \frac{28.02}{44.02} = 63.7\%$	$\text{O} = \frac{16.00}{44.02} = 36.3\%$	
6. NaOCl	$\text{Na} = \frac{22.99}{74.44} = 30.9\%$	$\text{O} = \frac{16.00}{74.44} = 21.5\%$	$\text{Cl} = \frac{35.45}{74.44} = 47.6\%$
7. Al_2S_3	$\text{Al} = \frac{53.94}{150.17} = 35.9\%$	$\text{S} = \frac{96.21}{150.17} = 64.1\%$	

Name _____

Naming Binary Compounds

Directions:

1. Using your periodic table, find the Oxidation Number for each element.
2. Write the elements. The (+) ion (cation) is written first.
3. "Criss-Cross-and-Reduce" the oxidation numbers so that they are now subscripts.
4. Do not write the number "1", it is assumed.
5. When naming your new compound, the "-" ion (anion) gets the "-ide" ending and is written second.



Figure 19.18: The criss-cross method is a simple way to determine the chemical formula of a compound.

Determine the formulas for the following:

	Ions	Symbols with Oxidation Numbers	Formula	Compound Name
1)	Lithium and bromine	Li^{+1} Br^{-1}	LiBr	Lithium Bromide
2)	Potassium and iodine			
3)	Hydrogen and phosphorus			
4)	Sodium and nitrogen			
5)	Magnesium and fluorine			
6)	Calcium and chlorine			
7)	Aluminum and oxygen			
8)	Boron and sulfur			

Write the oxidation numbers and name of the compounds below:

	Formula	Symbols and Oxidation Numbers	Compound Name
1)	MgBr_2	Mg^{+2} Br^{-1}	Magnesium Bromide
2)	BaF_2		
3)	BeBr_2		
4)	K_2S		
5)	Li_3N		
6)	Ca_3P_2		
7)	Na_2O		

*Extra Credit (+2 points): Look up one of the compounds on this page, give the formula, a brief description of its appearance, and one of its uses. Write your answer on the back of this worksheet.
Liz LaRosa 5th Grade Science www.middle-school-science.com 2009

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