

Binary ionic compounds names and formulas worksheet answers

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Naming Ionic Compounds Practice Worksheet

Name the following ionic compounds:

- 1) NH₄Cl _____
 2) Fe(NO₃)₃ _____
 3) TiBr₃ _____
 4) Cu₃P _____
 5) SnSe₂ _____
 6) GaAs _____
 7) Pb(SO₄)₂ _____
 8) Be(HCO₃)₂ _____
 9) Mn₂(SO₄)₃ _____
 10) Al(CN)₃ _____

Write the formulas for the following compounds:

- 11) chromium (VI) phosphate _____
 12) vanadium (IV) carbonate _____
 13) tin (II) nitrite _____
 14) cobalt (III) oxide _____
 15) titanium (II) acetate _____
 16) vanadium (V) sulfide _____
 17) chromium (III) hydroxide _____
 18) lithium iodide _____
 19) lead (II) nitride _____
 20) silver bromide _____

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Name Key
Period ___ Date ___/___/___**3 • Molecules and Compounds****MOLAR MASS & % COMPOSITION****I. Molar Masses**

Given a periodic table, you should be able to calculate the molecular mass (in u's) or the molar mass (in grams) for any element or compound.

Examples: (give answers to two decimal places)

H ₂ SO ₄	Cl ₂	Ca(OH) ₂	HC ₂ H ₃ O ₂
98.09 amu	70.90 amu	74.10 amu	60.06 amu
49.01 amu	35.45 amu	37.38 amu	30.03 amu

II. Fraction and Percent CompositionIt is useful to determine how much of a compound's mass is made up of each element. Water, H₂O, for example has a molar mass of 18.02 g. The H's mass is 2(1.0079) = 2.02 g. The O's mass is 16.00 g.We can set up fractions for each element: H = $\frac{2.02}{18.02} = 0.112 = 11.2\%$, O = $\frac{16.00}{18.02} = 0.888 = 88.8\%$.This is called the **percent composition**. The fraction composition is a good in-between step.

Determine the fraction and percent composition of each element below (answer to one decimal place):

1. H ₂ SO ₄	$\frac{2.02}{98.09} = 2.1\%$	$S = \frac{32.07}{98.09} = 32.7\%$	$O = \frac{64.00}{98.09} = 65.2\%$
2. Ca(OH) ₂	$\frac{40.08}{74.10} = 54.1\%$	$O = \frac{32.00}{74.10} = 43.2\%$	$H = \frac{2.02}{74.10} = 2.7\%$
3. HC ₂ H ₃ O ₂	$\frac{12.01}{60.06} = 20.0\%$	$C = \frac{24.02}{60.06} = 40.0\%$	$O = \frac{32.00}{60.06} = 53.3\%$
4. CO ₂	$\frac{12.01}{44.01} = 27.3\%$	$O = \frac{32.00}{44.01} = 72.7\%$	
5. N ₂ O	$\frac{28.02}{44.01} = 63.7\%$	$O = \frac{16.00}{44.01} = 36.3\%$	
6. NaOCl	$\frac{22.99}{74.44} = 30.9\%$	$O = \frac{16.00}{74.44} = 21.5\%$	$Cl = \frac{35.45}{74.44} = 47.6\%$
7. Al ₂ S ₃	$\frac{53.98}{150.17} = 35.9\%$	$S = \frac{96.21}{150.17} = 64.1\%$	

Name _____

Naming Binary Compounds

Directions:

- Using your periodic table, find the Oxidation Number for each element.
- Write the elements. The (+) ion (cation) is written first.
- "Criss-Cross-and-Reduce" the oxidation numbers so that they are now subscripts.
- Do not write the number "1", it is assumed.
- When naming your new compound, the "-" ion (anion) gets the "-ide" ending and is written second.



Figure 19.18: The criss-cross method is a simple way to determine the chemical formula of a compound.

Determine the formulas for the following:

	Ions	Symbols with Oxidation Numbers	Formula	Compound Name
1) Lithium and bromine	Li ⁺¹ Br ⁻¹		LiBr	Lithium Bromide
2) Potassium and iodine				
3) Hydrogen and phosphorus				
4) Sodium and nitrogen				
5) Magnesium and fluorine				
6) Calcium and chlorine				
7) Aluminum and oxygen				
8) Boron and sulfur				

Write the oxidation numbers and name of the compounds below:

	Formula	Symbols and Oxidation Numbers	Compound Name
1) MgBr ₂	Mg ⁺² Br ⁻¹		Magnesium Bromide
2) BaF ₂			
3) BeBr ₂			
4) K ₂ S			
5) Li ₂ N			
6) Ca ₃ P ₂			
7) Na ₂ O			

*Extra Credit (+2 points): Look up one of the compounds on this page, give the formula, a brief description of its appearance, and one of its uses. Write your answer on the back of this worksheet.

Liz LaRosa 5th Grade Science www.middleschoolscience.com 2009

Naming Compounds Containing Polyatomic Ions

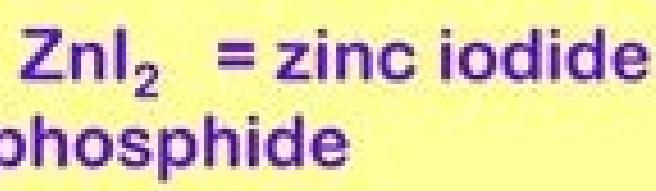
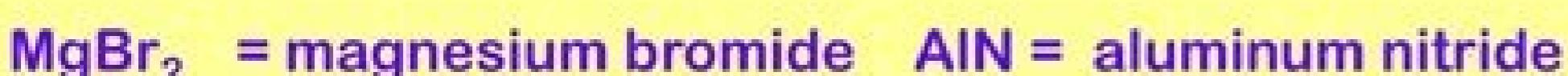
Use the following list of polyatomic ions to name the compounds given.

OH	hydroxide	ClO_3	chlorate	SO_3	sulfite
NO_3	nitrate	HCO_3	bicarbonate	PO_4	phosphate
$\text{C}_2\text{H}_3\text{O}_2$	acetate	SO_4	sulfate	NH_4	ammonium
		CO_3	carbonate		

- | | |
|--|---|
| 1. CuSO ₄ , _____ | 26. NaOH, _____ |
| 2. KClO ₃ , _____ | 27. NaNO ₃ , _____ |
| 3. NH ₄ OH, _____ | 28. Ba(OH) ₂ , _____ |
| 4. K ₂ CO ₃ , _____ | 29. K ₂ CO ₃ , _____ |
| 5. Na ₂ SO ₄ , _____ | 30. LiOH, _____ |
| 6. KC ₂ H ₃ O ₂ , _____ | 31. KHCO ₃ , _____ |
| 7. NH ₄ Cl, _____ | 32. NH ₄ F, _____ |
| 8. Mg(OH) ₂ , _____ | 33. KClO ₃ , _____ |
| 9. Zn(NO ₃) ₂ , _____ | 34. NaC ₂ H ₃ O ₂ , _____ |
| 10. MgSO ₄ , _____ | 35. CaClO ₂ , _____ |
| 11. CaCO ₃ , _____ | 36. SrNO ₃ , _____ |
| 12. KOH, _____ | 37. Be(OH) ₂ , _____ |
| 13. Ca(OH) ₂ , _____ | 38. Li ₂ SO ₄ , _____ |
| 14. K ₂ CO ₃ , _____ | 39. Al ₂ (SO ₄) ₃ , _____ |
| 15. AlPO ₄ , _____ | 40. BaSO ₄ , _____ |
| 16. Na ₃ PO ₄ , _____ | 41. NH ₄ I, _____ |
| 17. MgSO ₄ , _____ | 42. NH ₄ HCO ₃ , _____ |
| 18. AgNO ₃ , _____ | 43. Cu ₃ (PO ₄) ₂ , _____ |
| 19. Na ₂ CO ₃ , _____ | 44. Ni(OH) ₂ , _____ |
| 20. NaHCO ₃ , _____ | 45. ZnSO ₄ , _____ |
| 21. Ca(NO ₃) ₂ , _____ | 46. MnCO ₃ , _____ |
| 22. BaSO ₄ , _____ | 47. HgOH, _____ |
| 23. BPO ₄ , _____ | 48. Au ₂ SO ₄ , _____ |
| 24. LiNO ₃ , _____ | 49. NH ₄ NO ₃ , _____ |
| 25. Al(NO ₃) ₃ , _____ | 50. LiClO ₃ , _____ |

Nomenclature of binary ionic compounds

Answers:



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